

BP-102T

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ODD SEMESTER EXAMINATION , 2022-23

**COURSE NAME :- B.PHARM**

**SEMESTER- I<sup>ST</sup>**

**SUBJECT :- PHARMACEUTICAL ANALYSIS**

TIME: 3 HOURS

MAX MARKS:75

NOTE: Attempt all parts.

**(1x20)**

### **PART A**

1. The number of gms of solute per 100 ml of solvent is known as

- a) Normality                      b) % weight by volume                      c) Molality                      d) Mole fraction

2. The number of gm-equivalent of the solute per liter of the solution is known as

- a) Normality                      b) Molarity                      c) Molality                      d) Mole fraction

3. Which of the following is a general property of bases?

- a) Taste sour    b) Turn litmus paper to red    c) Conduct electric current in solution    d) None

4. A buffer solution may contain equal moles of

- a) Weak acid and strong base                      b) Strong acid and strong base  
c) Weak acid and its conjugated base                      d) Strong acid and its conjugated base

5. In titration end point can be determined by change in colour by

- a) Measuring cylinder                      b) Burette                      c) Instruments                      d) Indicator

6. The quantity of chemical in each liter of solution is known as

- a) Normality                      b) Strength                      c) Molecular weight                      d) Equivalent weight

7. An example of a primary standard solution is

- a) FeSO<sub>4</sub>                      b) Na<sub>2</sub>CO<sub>3</sub>                      c) NH<sub>4</sub>OH                      d) NaOH

8. A solution of known concentration is Known as

- a) Buffer solution                      b) Neutral solution                      c) Standard solution                      d) Saturated solution

9. An Arrhenius acid is defined as a chemical species that

- a) is a proton donor b) is a proton acceptor c) Produces  $H^+$  ion in solution d) Produces  $OH^-$  ion
10. The pH at which an indicator changes colour is known as
- a) Standard point b) Transition point c) Equivalence point d) None of these
11. Non aqueous titration is carried out for
- a) Water insoluble drugs b) Weakly acidic drugs c) Weakly basic drugs d) all the above
12. Which one is aprotic solvent?
- a) Chloroform b) Benzene c) Both d) None
13. Indicator used in Complexometric titration is.....
- a) Erichrome black T b) Xylenolorange c) Mordant black II d) All
14. SI unit of conductance is..
- a) Mho b) Seimens c) Volt d) None of these
15. Potentiometry is an.....method of analysis
- a) Spectroscopic b) Electrometric c) Analytical d) None of these
16. In a redox reaction ,
- a) Oxidation occurs b) Reduction occurs c) both d) None of these
17. Perchloric acid can be standardized by using
- a) Benzoic acid b) Oxalic acid c) Potassium hydrogen phthalate d) Tartaric acid
18. Which one is used as indicator for non aqueous titration.
- a) Crystal violet b) Thymol blue c) Oracet Blue B d) All
19. Titration based on the use of silver nitrate are called..... Titration
- a) Argentometric b) Complexometric c) Amperometric d) conductometric
20. The end point for an EDTA titration is usually found by using a.....indicator
- a) Metallochromic b) Redox c) Acid base d) All

## PART B

(ATTEMPT ANY 2)

(2x10)

1. Describe the principle and procedure involved in the preparation and standardization of ceric ammonium sulphate solution.
2. Interpret Fajan's method in precipitation titration.
3. Explain impurities and its sources of in details.

## PART C

(ATTEMPT ANY 7)

(7x5)

1. Describe the principle and procedure involved in the limit test for chloride.
2. Explain Errors and define the types of Errors.
3. Interpret the standardization method of 0.1N AgNO<sub>3</sub> solution.
4. Define the theories of acid base titration.
5. Explain different types of Complexometric titration.
6. Summarize short notes on Co-Precipitation.
7. Explain the principle and theory involved in the conductometry.
8. Describe an indicator used in redox titration.
9. Explain the working principle of saturated calomel electrode.